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STA. STUDY.

PLAY. Solute. what

is dissolved.

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Solvent. what does  
the dissolving.

Soluble. can be  
dissolved. Tyndall

Effect. dispersal of  
particles; ex. seeing  
beams of light

through fog  
(scattering light)

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Capillary Action.

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Tension.

Vaporization. the

attraction of the

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to the surface of  
a.... a substance  
that can flow and  
therefore take the  
shape of its.... a  
force that tends to  
pull adjacent parts  
of a liquid's  
surfac....

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Mixtures. Solutions are homogeneous mixtures of two or more substances in a single phase.

Soluble describes a substance as capable of being dissolved. Solvent

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is the dissolving medium in a solution. Solute is the substance that is dissolved in a

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of Mixtures The  
Solution Process

Concentrations of  
Solution 1 2 Section

12.1 Types of  
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401-406 2 3

Vocabulary Colloid  
Dispersed phase  
Dispersing medium

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Electrolyte

Nonelectrolyte

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Nonelectrolyte.

Solute. Solvent.

Solution.



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homogeneous  
mixture of two or  
more substances in  
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Teacher Notes and  
Answers Chapter

12 SECTION 1

SOLUTIONS Mixed  
SHORT ANSWER

1. c 2. a 3. b 2. a.

alcohol b. water c.

the gels 3. The

mixture is a colloid.

The properties are

consistent with

those reported in

Table 3 on page

404 of the text. The

particle size is

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small, but not too  
small, and the  
mixture

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In the space

provided, write the  
letter of the term or  
phrase that best  
completes each  
statement or best



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Answers each  
question. \_\_\_\_\_ 1.  
Agitation prevents  
settling in a(n) a.  
alloy. b.

homogeneous  
mixture. c.  
suspension. d.  
gaseous mixture.

\_\_\_\_\_ 2. All of the  
following are  
heterogeneous ...

Assessment

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QUESTIONS. 13.1.

Classify the  
following amines as  
primary, secondary  
and tertiary: Ans:

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(i)  $1^\circ$  (ii)  $3^\circ$  (iii)

$1^\circ$  (iv)  $2^\circ$  13.2.

Write the structures of different isomeric amines

corresponding to the molecular formula,  $C_4H_{11}N$ . (i) Write the IUPAC names of all the isomers

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13 Amines

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3.3: Determine the

amount in moles of

the following:a.

12.15 g Mg... Get

solutions

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Nr. 64.

ladkowska, J.

Polynômes quasi-  
univalents et  
univalents. 1960.

This book had its  
nucleus in some

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lectures given by  
one of us (J. O'M. B.  
) in a course on  
electrochemistry to  
students of energy  
conversion at the  
University of  
Pennsylvania. It  
was there that he  
met a number of  
people trained in  
chemistry, physics,  
biology, metallurgy,  
and materials



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wanted to know  
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something about  
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electrochemistry.

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The concept of  
writing a book  
about  
electrochemistry  
wh ich could be  
understood by  
people with very  
varied backgrounds  
was thereby  
engendered. The

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lectures were recorded and written up by Dr. Klaus Muller as a 293-page manuscript. At a later stage, A. K. N. R. joined the effort; it was decided to make a fresh start and to write a much more comprehensive text. Of methods for

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direct energy conversion, the electrochemical one is the most advanced and seems the most likely to become of considerable practical importanee. Thus, conversion to electrochemically powered trans portation systems

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appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met. Corrosion is recognized as having an electrochemical

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basis. The synthesis of nylon now contains an important

electrochemical stage. Some central biological mechanisms have been shown to take place by means of electrochemical reactions. A number of American organizations have

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recently recommended

greatly increased

activity in training

and research in

electrochemistry at

universities in the

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the enormous

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Chemistry, farmers,  
and research  
personnel

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industry. It covers:

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Laboratory methods  
of analysis The

book also includes

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