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*SOLUTIONS/ANSWERED IB Chemistry SL/HL Topic 1: Pearson (2014) Textbook Practice Questions*

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$4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightleftharpoons 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$   $\Delta H^\ominus = -910 \text{ kJ mol}^{-1}$

Describe the effect, if any, of each of the following changes on the equilibrium concentration of nitrogen monoxide in a particular equilibrium mixture, giving a reason in each case. (i) Increasing the pressure, at constant temperature [2]

### **CHEMISTRY STANDARD LEVEL PAPER 2**

– 2 – Section A Answer all questions. Write your answers in the boxes provided. 1.

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Ethanedioic acid is a diprotic acid. A student determined the value of  $x$  in the formula of hydrated ethanedioic acid,  $\text{HOOC-COOH} \cdot x\text{H}_2\text{O}$ , by titrating a known mass of the acid with a  $0.100 \text{ mol dm}^{-3}$  solution of  $\text{NaOH (aq)}$ .  $0.795 \text{ g}$  of ethanedioic acid was dissolved in distilled water and made up to a total ...

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Ethane,  $\text{C}_2\text{H}_6$ , and disilane,  $\text{Si}_2\text{H}_6$ , are both hydrides of group 4 elements with similar structures but with different chemical

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properties. (a) (i) Deduce the Lewis (electron dot) structure for  $\text{Si}_2\text{H}_6$  showing all valence electrons. [1] (ii) State and explain the H–Si–H bond angle in  $\text{Si}_2\text{H}_6$ . [2]

## **CHEMISTRY Candidate session number STANDARD LEVEL PAPER 2**

[2] (b) Marit arranged the values she found in Table 12 into an energy cycle.

$\text{C}_2\text{H}_4(\text{g})$	+	$\text{H}_2(\text{g})$		$\text{C}_2\text{H}_6(\text{g})$	3 0 2		$2\text{CO}_2(\text{g})$	+	$3\text{H}_2\text{O}$	– 1 4 1 1
kJ mol <sup>-1</sup> – 1 – 2 8 6 kJ mol <sup>-1</sup> – 1 – 1 5 6 0										

kJ mol<sup>-1</sup> (hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....



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