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$M_{\text{nitrogen}} + 0.21 M_{\text{oxygen}} =$

$0.79(28) + 0.21(32) = 28.84$  The

mass fractions of oxygen and  
nitrogen in air are then

$m_{\text{oxygen}} = (0.21)(32)/28.84 =$

$0.233$ , or  $23.3\%$  and  $m_{\text{nitrogen}} =$

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4. Write additional data required to solve the problem and the chemical equations if the process involves chemical reaction. 5. Select a suitable basis of calculations. 6. List by symbols each of the unknown values of

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the stream flows and  
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balances being the trunk and the subjects of thermodynamics, fluid flow, heat transfer, mass transfer, reactor kinetics, process control, and process design being the branches off the trunk. From this perspective, it is



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in Chemical Engineering

$z + \Delta z$   $P(z) = \rho g z$  (2.2.8) Setting the  
limit on the left hand side as  $\Delta z \rightarrow 0$

gives the definition of the

derivative of  $P$  with respect to  $z$ :

$\frac{dP}{dz} = \rho g$  (2.2.9) In other words,

the change in pressure,  $P$ , with

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height,  $z$ , is negative – pressure reduces with height (or increases with depth – as any diver knows).

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