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Chemical equation for the reaction at equilibrium is written using double arrows to indicate the overall reversibility of the reaction.

$$2\text{HgO}(s) \rightleftharpoons 2\text{Hg}(l) + \text{O}_2(g)$$

Equilibrium, a Dynamic State Many chemical

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reactions are reversible
under ordinary conditions of
temperature and
concentration. They will
reach a state of equilibrium

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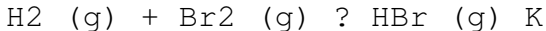
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$\approx 4.0 \times 10^{-2}$ $K = 3.5 \times 10^4$

Chapter 18 Chemical
Equilibrium 18.2 Shifting
Equilibrium The Effect of
Change in Pressure: Ways to
Change Pressure Add or
remove a gaseous reactant or
product Add an inert gas

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(one not involved in the reaction) An inert gas increases the total pressure but has no effect on the concentrations or partial pressures of the reactants or products

The Effect of Change in Pressure

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the volume of the container
When the volume of the ...

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Equilibrium SECTION 2 SHORT

ANSWER Answer the following

questions in the space

provided. 1. _____ Raising

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Answers Section 3
the temperature of any
equilibrium system always
(a) favors the forward
reaction. (b) favors the
reverse reaction. (c) favors
the exothermic reaction.

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Equilibrium Answers

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Equilibrium SECTION 2 SHORT

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Equilibrium Answers Section

1

562 Chapter 18 Chemical
Equilibrium Figure 18-3 If
the only way in or out of
San Francisco and Sausalito
is across the Golden Gate
Bridge, which joins the two

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Answers Section 3
cities, then the number of vehicles in the two cities will remain constant if the number of vehicles per hour crossing the bridge in one direction equals the number of vehi-

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Chapter 18: Chemical Equilibrium

5. Amongst the following chemical reactions, the irreversible reaction is: a) $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$. b) $\text{H}_2 + \text{I}_2 \rightleftharpoons \text{HI}$. c) $\text{CaCO}_3 \rightleftharpoons \text{CaO} + \text{CO}_2$. d) O_2

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Answers Section 3. ANS: The reaction is not reversible.

6. When the rate of forward reaction becomes equal to backward reaction, this state is termed as: a) Chemical equilibrium. b) Reversible state

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